**AP Chem Chp3 Stoichiometry**

Wkst: Molar mass, % Composition

Show work!! Use sig figs to round your answers.

**Part 1: Find the molar mass for each.**

1. Aluminum nitrate

2. Iron (III) hydroxide

**Part 2: Find the number of grams for each sample listed**

3. 0.30 mol NaHCO3

4. 1.50 mol H2SO4

5. 1.26 x 10-4 mol sodium dihydrogen phosphate

**Part 3: Calculate the number of moles for each sample listed**

6. 17.1 g H2S

7. 10.08 g calcium nitrate

8. 8.8 g potassium carbonate

**Part 4: Find the mass in grams for each quantity.**

9. 1.788 x 1029 atoms of carbon

10. 9.4044 x 1015 molecules of water.

**Part 5: Identify the type of representative particle for each sample listed and then find the number of particles for each quantity**.

11. 19.30 g isopentyl acetate (C7H14O2)

12. 290.45 g copper (II) nitrate

**Part 6: Find the percentage composition.**

13. Calcium phosphate

14. A sample of benzene is analyzed and found to consist of 13.74 g carbon and 1.15 g of hydrogen. Find the percent composition.