

Part A: Naming Ionic Compounds
(Name the Following Ionic Compounds.)

1. NH_4Cl _____
2. $\text{Fe}(\text{NO}_3)_3$ _____
3. CaBr_2 _____
4. Cu_3P _____
5. SnSe_2 _____
6. GaS _____
7. $\text{Pb}(\text{SO}_4)_2$ _____
8. $\text{Be}(\text{HCO}_3)_2$ _____
9. $\text{Al}_2(\text{SO}_3)_3$ _____
10. $\text{Ni}(\text{CN})_2$ _____

Part B: Formula Writing of Ionic Compounds
(Write Formulas for the Following Ionic Compounds.)

1. chromium (III) phosphate _____
2. tin (II) nitrite _____
3. silver bromide _____
4. copper (II) sulfate _____
5. magnesium acetate _____
6. lead (II) nitride _____
7. strontium nitride _____
8. lithium iodide _____
9. copper (I) oxide _____
10. cobalt (III) oxide _____