

## WORKSHEET 7.A

NAME Key

## IONIC BINARY COMPOUNDS

DIRECTIONS: Using the ion charge numbers, write the formula and name of the compound formed by the two elements. Include the ion charges as part of the formulas.

- Sodium<sup>+1</sup> and Sulfur<sup>-2</sup>  
Sodium Sulfide  $\text{Na}_2\text{S}$
- Barium<sup>+2</sup> and oxygen<sup>-2</sup>  
Barium Oxide  $\text{BaO}$
- Rubidium<sup>+1</sup> and tellurium<sup>-2</sup>  
Rubidium Telluride  $\text{Rb}_2\text{Te}$
- Lithium<sup>+1</sup> and oxygen<sup>-2</sup>  
Lithium Oxide  $\text{Li}_2\text{O}$
- Francium<sup>+1</sup> and sulfur<sup>-2</sup>  
Francium Sulfide  $\text{Fr}_2\text{S}$
- Potassium<sup>+1</sup> and selenium<sup>-2</sup>  
Potassium Selenide  $\text{K}_2\text{Se}$
- Lithium<sup>+1</sup> and chlorine<sup>-1</sup>  
Lithium Chloride  $\text{LiCl}$
- Barium<sup>+2</sup> and sulfur<sup>-2</sup>  
Barium Sulfide  $\text{BaS}$
- Aluminum<sup>+3</sup> and nitrogen<sup>-3</sup>  
Aluminum Nitride  $\text{AlN}$
- Strontium<sup>+2</sup> and fluorine<sup>-1</sup>  
Strontium Fluoride  $\text{SrF}_2$
- Barium<sup>+2</sup> and bromine<sup>-1</sup>  
Barium Bromide  $\text{BaBr}_2$
- Aluminum<sup>+3</sup> and oxygen<sup>-2</sup>  
Aluminum Oxide  $\text{Al}_2\text{O}_3$
- Beryllium<sup>+2</sup> and iodine<sup>-1</sup>  
Beryllium Iodide  $\text{BeI}_2$
- Cesium<sup>+1</sup> and astatine<sup>-1</sup>  
Cesium astatide  $\text{CsAt}$
- Magnesium<sup>+2</sup> and fluorine<sup>-1</sup>  
magnesium fluoride  $\text{MgF}_2$
- Radium<sup>+2</sup> and bromine<sup>-1</sup>  
Radium Bromide  $\text{RaBr}_2$
- Calcium<sup>+2</sup> and chlorine<sup>-1</sup>  
Calcium Chloride  $\text{CaCl}_2$
- Radium<sup>+2</sup> and bromine<sup>-1</sup>  
Radium Bromide  $\text{RaBr}_2$
- Potassium<sup>+1</sup> and iodine<sup>-1</sup>  
Potassium Iodide  $\text{KI}$
- Sodium<sup>+1</sup> and chlorine<sup>-1</sup>  
Sodium Chloride  $\text{NaCl}$
- Magnesium<sup>+2</sup> and sulfur<sup>-2</sup>  
Magnesium Sulfide  $\text{MgS}$
- Calcium<sup>+2</sup> and chlorine<sup>-1</sup>  
Calcium Chloride  $\text{CaCl}_2$
- Aluminum<sup>+3</sup> and bromine<sup>-1</sup>  
Aluminum Bromide  $\text{AlBr}_3$
- Radium<sup>+2</sup> and iodine<sup>-1</sup>  
Radium Iodide  $\text{RaI}_2$

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- Zinc<sup>+2</sup> and oxygen<sup>+2</sup>  
Zinc Oxide  $ZnO$
- Hydrogen<sup>+1</sup> and fluorine<sup>-1</sup>  
Hydrogen Fluoride  $HF$
- Iron<sup>+2</sup>(II) and bromine<sup>-1</sup>  
Iron II Bromide  $FeBr_2$
- Nickel<sup>+2</sup>(II) and oxygen<sup>-2</sup>  
Nickel II Oxide  $NiO$
- Copper<sup>+2</sup>(II) and iodine<sup>-1</sup>  
Copper II Iodide  $CuI_2$
- Iron<sup>+2</sup>(II) and oxygen<sup>-2</sup>  
Iron II Oxide  $FeO$
- Mercury<sup>+2</sup>(II) and fluorine<sup>-1</sup>  
Mercury II Fluoride  $HgF_2$
- Barium<sup>+2</sup> and sulfur<sup>-2</sup>  
Barium Sulfide  $BaS$
- Iron<sup>+2</sup>(II) and oxygen<sup>-2</sup>  
Iron II Oxide  $FeO$
- Strontium<sup>+2</sup> and oxygen<sup>-2</sup>  
Strontium Oxide  $SrO$
- Aluminum<sup>+3</sup> and sulfur<sup>-2</sup>  
Aluminum Sulfide  $Al_2S_3$
- Copper<sup>+1</sup>(I) and nitrogen<sup>-3</sup>  
Copper I Nitride  $Cu_3N$
- Silver<sup>+1</sup> and sulfur<sup>-2</sup>  
Silver Sulfide  $Ag_2S$
- Iron<sup>+3</sup>(III) and bromine<sup>-1</sup>  
Iron III bromide  $FeBr_3$
- Cobalt<sup>+2</sup>(II) and chlorine<sup>-1</sup>  
Cobalt II Chloride  $CoCl_2$
- Copper<sup>+1</sup>(I) and oxygen<sup>-2</sup>  
Copper I Oxide  $Cu_2O$
- Chromium<sup>+2</sup>(II) and bromine<sup>-1</sup>  
Chromium II bromide  $CrBr_2$
- Tin<sup>+4</sup>(IV) and sulfur<sup>-2</sup>  
Tin IV Sulfide  $SnS_2$
- Lead<sup>+2</sup>(II) and oxygen<sup>-2</sup>  
Lead II Oxide  $PbO$
- Iron<sup>+3</sup>(III) and sulfur<sup>-2</sup>  
Iron III Sulfide  $Fe_2S_3$
- Lead<sup>+4</sup>(IV) and iodine<sup>-1</sup>  
Lead IV Iodide  $PbI_4$
- Calcium<sup>+2</sup> and bromine<sup>-1</sup>  
Calcium Bromide  $CaBr_2$
- Tin<sup>+2</sup>(II) and nitrogen<sup>-3</sup>  
Tin II nitride  $Sn_3N_2$
- Mercury<sup>+2</sup>(II) and iodine<sup>-1</sup>  
Mercury II Iodide  $HgI_2$