

Chemistry Ch 14
Worksheet 14B
Molarity

Fill in the blank spaces in the table below. Show all of your work when solving the problems.

Compounds used are:

1. Calcium Chloride
2. Iron III nitrate
3. Sodium carbonate
4. Magnesium Sulfate

Formula of solute	Grams Dissolved	Moles Dissolved	Volume of solution	Molarity
1.	15.00 g		0.200 L	
2.			1.00 L	1.00 M
3.		0.035 moles		1.50 M
4.			300. mL	0.024 M

5. Describe how you would prepare 1.50 L of 0.25 M solution of sodium sulfate
6. How many grams of ammonium carbonate are needed to make 3.50 L of a 1.55 M solution?
7. How many grams of calcium carbonate are in 2.50 L of a 0.25 M solution?
8. How many moles of zinc sulfate are contained in 4.33 L of a 0.75 M solution?
9. Calculate the number of silver nitrate formula units contained in 2.00 L of a 0.55 M solution. (hint: Find moles first, then formula units)
10. A student needs to prepare 200. mL of a 1.25 M ammonium chloride solution. Describe the steps for the student to use.