

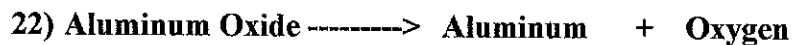
Matching

- | | |
|---------------------------------------|--|
| _____ 1) Reactants | A) (s) |
| _____ 2) Exothermic Reaction | B) A substance that speeds up a reaction |
| _____ 3) Double Replacement Reaction | C) yields or produces |
| _____ 4) Catalyst | D) Mass of reactants equals mass of products |
| _____ 5) Endothermic Reaction | E) (l) |
| _____ 6) Precipitate | F) (aq) Dissolved in water |
| _____ 7) -----> | G) Equation with the same number of atoms on both sides of the equation |
| _____ 8) Coefficient | H) A solid produced in a chemical reaction |
| _____ 9) Law of Conservation of Mass | I) A substance that slows down or blocks a reaction |
| _____ 10) Balanced Equation | J) A reaction that takes in heat to break bonds (feels cold) |
| _____ 11) Decomposition Reaction | K) A reaction that gives off heat or light |
| _____ 12) Single Replacement Reaction | L) A reaction in which two or more substances combine to form a new substance |
| _____ 13) Products | M) The number of units of each substance involved in a reaction |
| _____ 14) Diatomic elements | N) (g) |
| _____ 15) Inhibitor | O) Elements like hydrogen and oxygen when they are not in a compound |
| _____ 16) Solid | P) Substances on the left of the reaction |
| _____ 17) Liquid | Q) $AB + CD \text{-----} \rightarrow AD + CB$ |
| _____ 18) Synthesis Reaction | R) Substances on the right of the reaction |
| _____ 19) Aqueous Solution | S) $A + BC \text{-----} \rightarrow AC + B$ |
| _____ 20) Gas | T) A reaction in which a substance is broken into two or more substances (requires a catalyst) |

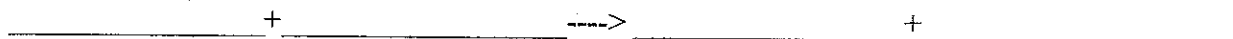
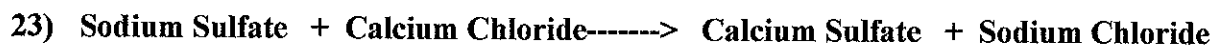
Write the balanced equations below and identify the type of reaction they represent. (3pts/each)



Type of reaction: _____



Type of reaction: _____



Type of reaction: _____



Type of reaction: _____



Type of reaction: _____