

PS Ch20

WORKSHEET Naming & Formulas

NAME _____

IONIC BINARY COMPOUNDS

DIRECTIONS: Using the ion charge numbers, write the formula and name of the compound formed by the two elements. Include the ion charges as part of the formulas.

1. Sodium and Sulfur

2. Barium and oxygen

3. Rubidium and tellurium ^{Te⁻²}

4. Lithium and oxygen

5. Francium and sulfur

6. Potassium and selenium ^{Se⁻²}

7. Lithium and chlorine

8. Barium and sulfur

9. Aluminum and nitrogen

10. Strontium and fluorine

11. Barium and bromine

12. Aluminum and oxygen

13. Beryllium and iodine ^{Be⁺²}

14. Cesium and astatine ^{At}

15. Magnesium and fluorine

16. Radium and bromine

17. Calcium and chlorine

18. Radium and bromine

19. Potassium and iodine

20. Sodium and chlorine

21. Magnesium and sulfur

22. Calcium and chlorine

23. Aluminum and bromine

24. Radium and iodine

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IONIC BINARY COMPOUNDS

DIRECTIONS: Using the ion charge numbers, write the formula and name of the compound formed by the two elements. Include the ion charges as part of the formulas.

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|--|------------------------------|
| 1. Zinc and oxygen | 13. Silver and sulfur |
| 2. Hydrogen and fluorine | 14. Iron(III) and bromine |
| 3. Iron(II) and bromine | 15. Cobalt(II) and chlorine |
| 4. Nickel(II) and oxygen | 16. Copper(I) and oxygen |
| 5. Copper(II) and iodine | 17. Chromium(II) and bromine |
| 6. Iron(II) and oxygen | 18. Tin(IV) and sulfur |
| 7. Mercury(II) and fluorine | 19. Lead(II) and oxygen |
| 8. Barium and sulfur | 20. Iron(III) and sulfur |
| 9. Iron(II) and oxygen | 21. Lead(IV) and iodine |
| 10. ^{Sr} Strontium and oxygen | 22. Calcium and bromine |
| 11. Aluminum and sulfur | 23. Tin(II) and nitrogen |
| 12. Copper(I) and nitrogen | 24. Mercury(II) and iodine |