orksheet: Ionic Bonding, Naming Binary Compounds Period	Chapter 🗷 Cher	nical Bonds	Review 5	•	Name	
Frameworks Code) Part A — Ionic Bonding: Complete the following questions in the space provided: What is an ion? What is an ionic bond? How does an atom become a positive ion? How does an atom become a negative ion? How does an atom become a negative ion? How does the electrically neutral compounds? Give an example? How does the electrical conductivity of ionic compounds change when they are melted or dissolved in water? Why is this so? Part B - Naming binary compounds: Using the rules found on page 117 of your textbook for binary compounds, name the following: Name 14. Sn ₃ P ₄ 15. Cr ₂ S ₃ 16. Na ₂ S 11. BaF ₂ 17. AgCl 12. AlCl ₃ 18. FeBr ₃ 19. Cu ₂ O Part C: Use the chart answer the following questions Ions and Their Charges Name Charge Symbol/Formula Ammonium Potassium 1+ K* Oxide 2- O ² Calcium 2+ Mag ² + Phosphate 3- PO ₄ 2- Divide Symbol/Formula Magnesium 2+ Mag ² + Phosphate 3- PO ₄ Pol- Calcium 2+ Mag ² + Phosphate 3- PO ₄ 2- Divide Capplain.				ounds	Period	Date
What is an ion? What is an ionic bond? How does an atom become a positive ion? How does an atom become a negative ion? How do ions form electrically neutral compounds? Give an example? What characteristics do solid ionic compounds share? How does the electrical conductivity of ionic compounds change when they are melted or dissolved in water? Why is this so? **Art.B - Naming binary compounds: Using the rules found on page 117 of your textbook for binary compounds, name the following: **ABR** 14. Sn ₃ P ₄ 15. Cr ₂ S ₃ 10. AlN 16. Na ₂ S 11. BaF ₂ 17. AgCl 18. FeBr ₃ 19. Cu ₂ O **Part C: Use the chart answer the following questions **Ions and Their Charges** Name Charge Symbol/Formula Name Charge Symbol/Formula Ammonium 1+ NFI ₄ + Chloride 1- Cl Potassium 1+ K* Oxide 2- O ² Calcium 2+ Ca ²⁺ Sulfide 2- S ² Magnesium 2+ Mg ²⁺ Phosphate 3- PO ₄ ³⁻ 20. How many potassium ions are needed to balance the charge of one sulfide ion? Explain.						•
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What is an ionic bond? How does an atom become a positive ion? How does an atom become a negative ion? What characteristics do solid ionic compounds? Give an example? How does the electrical conductivity of ionic compounds change when they are melted or dissolved in water? Why is this so? Art B - Naming binary compounds: Using the rules found on page 117 of your textbook for binary compounds, name the following: NaBr		•				
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How does an atom become a negative ion? How do ions form electrically neutral compounds? Give an example? What characteristics do solid ionic compounds share? How does the electrical conductivity of ionic compounds change when they are melted or dissolved in water? Why is this so? ***art*B - Naming binary compounds: Using the rules found on page 117 of your textbook for binary compounds, name the following: NaBr						
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How does the electrical conductivity of ionic compounds change when they are melted or dissolved in water? Why is this so?	. What charact		•			
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art B - Naming binary compounds: Using the rules found on page 117 of your textbook for binary compounds, name the following: NaBr						
compounds, name the following: NaBr	water: wily i	.s uns so:				
compounds, name the following: NaBr	AD Namine	v hinary com			7 of your text	book for binary
14. Sn ₃ P ₄			•			
KCl 15. Cr ₂ S ₃	· . : ****	-		14. Sn ₃ P ₄		·
16. Na ₂ S 1. BaF ₂ 2. AlCl ₃ 3. ZnS 19. Cu ₂ O Part C: Use the chart answer the following questions ons and Their Charges Name Charge Symbol/Formula Name Charge Symbol/Formula Ammonium 1+ NH ₄ ⁺ Chloride 1- Cl Potassium 1+ K ⁺ Oxide 2- Calcium 2+ Ca ²⁺ Magnesium 2+ Mg ² + Phosphate 3- PO ₄ ³⁻ PO ₄ ³⁻ 20. How many potassium ions are needed to balance the charge of one sulfide ion? Explain.	. Nadi,	- ~ i				•
17. AgCl						
2. AlCl ₃ 18. FeBr ₃ 3. ZnS 19. Cu ₂ O Part C: Use the chart answer the following questions Sons and Their Charges Ions and Their Charges Name Charge Symbol/Formula Ammonium 1+ NH ₄ + Chloride 1- Cl Potassium 1+ K+ Oxide 2- O ² - Calcium 2+ Ca ²⁺ Sulfide 2- S ² - Magnesium 2+ Mg ² + Phosphate 3- PO ₄ ³ - 20. How many potassium ions are needed to balance the charge of one sulfide ion? Explain.					•	
3. ZnS Part C: Use the chart answer the following questions Sons and Their Charges Name Charge Symbol/Formula Name Charge Symbol/Formula Ammonium 1+ NH ₄ ⁺ Chloride 1- Cl Potassium 1+ K ⁺ Oxide 2- Calcium 2+ Ca ²⁺ Sulfide 2- Magnesium 2+ Mg ² + Phosphate 3- PO ₄ ³⁻ 20. How many potassium ions are needed to balance the charge of one sulfide ion? Explain.			•			
Part C: Use the chart answer the following questions Ions and Their Charges	- 					
Name Charge Symbol/Formula Name Charge Symbol/Formula Ammonium 1+ NH ₄ ⁺ Chloride 1- Cl Potassium 1+ K ⁺ Oxide 2- O ² Calcium 2+ Ca ²⁺ Sulfide 2- S ² Magnesium 2+ Mg ² + Phosphate 3- PO ₄ ³⁻ 20. How many potassium ions are needed to balance the charge of one sulfide ion? Explain.				· -		
NameChargeSymbol/FormulaNameChargeSymbol/FormulaAmmonium $1+$ NH_4^+ Chloride $1 CI^-$ Potassium $1+$ K^+ Oxide $2 O^{2-}$ Calcium $2+$ Ca^{2+} Sulfide $2 S^{2-}$ Magnesium $2+$ Mg^2+ Phosphate $3 PO_4^{3-}$ 20. How many potassium ions are needed to balance the charge of one sulfide ion? Explain.			the following question		harges	-
Ammonium 1+ NH ₄ + Chloride 1- Cl Potassium 1+ K+ Oxide 2- O ² Calcium 2+ Ca ²⁺ Sulfide 2- S ²⁻ Magnesium 2+ Mg ² + Phosphate 3- PO ₄ ³⁻ 20. How many potassium ions are needed to balance the charge of one sulfide ion? Explain.		,	Symbol/Formula			Symbol/Formula
Potassium 1+ K ⁺ Oxide 2- O ²⁻ Calcium 2+ Ca ²⁺ Sulfide 2- S ²⁻ Magnesium 2+ Mg ² + Phosphate 3- PO ₄ ³⁻ 20. How many potassium ions are needed to balance the charge of one sulfide ion? Explain.						
Potassium Calcium 2+ Ca ²⁺ Sulfide 2- S ²⁻ Magnesium 2+ Mg ² + Phosphate 3- PO ₄ ³⁻ 20. How many potassium ions are needed to balance the charge of one sulfide ion? Explain.				<u></u>		l
Magnesium 2+ Ca Surface 2 Phosphate 3- PO ₄ ³ - 20. How many potassium ions are needed to balance the charge of one sulfide ion? Explain.						
20. How many potassium ions are needed to balance the charge of one sulfide ion? Explain.	Calcium					
		1 '	, -	-	1	<u>-</u>
	20. How many p	ootassium ion	s are needed to balance	the charge of one s	ulfide ion? Ex	cpiain.
21. Predict the formula for calcium chloride:, & potassium phosphate						
	22. Which ions	in the table ar	e polyatomic ions?			