

## Ch 11 Review

Write a balanced equation, not the given and unknowns.

1. Define stoichiometry
  
2. 25.9g of Phosphoric acid ( $\text{H}_3\text{PO}_4$ ) reacts completely with calcium hydroxide.
  - a. How many grams of water are produced?
  - b. What type of reaction is this?
  - c. How many moles of calcium hydroxide are required?
  - d. How many grams of the other product (not water) will be produced?
  
3. Aluminum sulfide forms from the reaction of 9.00g of Aluminum and 10.0 g of sulfur.
  - a. Which reactant is the limiting factor?
  - b. How many grams of excess reactant will be left?
  - c. How many grams of Aluminum sulfide will be produced?
  
4. 24.5g of  $\text{C}_3\text{H}_8$  are combusted with 45.0g of Oxygen.
  - a. How many moles of each reactant are needed?
  - b. Which reactant is the limiting factor?
  - c. How many grams of each product are produced?