

AP Chem Unit 1 Ch2 Handout

- ① Law of Multi Proportions Ratios of the masses of N combining w/ 1g of O in each pair of compounds should be small whole #'s

$$\frac{A}{B} = \frac{1.7500g N}{.8750g N} = \frac{2}{1}$$

$$\frac{B}{C} = \frac{.8750g N}{.4375g N} = \frac{2}{1}$$

$$\frac{A}{C} = \frac{1.7500g N}{.4375g N} = \frac{4}{1}$$

These Results Support the Law of Multi proportions

- ② ${}_{9}^{19}\text{F}$ 9 electrons
10 Neutrons

- ③ A) Cesium Fluoride
B) Aluminum Chloride
C) Lithium Hydride

- ④ A) Copper I chloride D) Manganese IV oxide
B) Mercury II oxide e) Lead II Chloride
C) Iron III oxide

- ⑤ A) Cobalt II Bromide
B) Calcium Chloride
C) Aluminum Oxide
D) Chromium III Chloride

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- 6) A) Sodium Sulfate
B) Potassium dihydrogen
C) Iron III Nitrate
D) Manganese II hydroxide
E) Sodium Sulfite
F) Sodium Carbonate
G) Sodium hydro Carbonate
H) Cesium Perchlorate
I) Sodium hydrochlorite
J) Sodium Selenate
K) Potassium Bromate

- 7) A) Phosphorus pentachloride
B) Phosphorus trichloride
C) Sulfur trioxide
D) Sulfur hexafluoride
E) Sulfur dioxide
F) Carbon dioxide

- 8) A) HBr hydrobromic Acid
B) HBrO hydrobromous Acid
C) HBrO₂ Bromous Acid
D) HBrO₃ Bromic Acid
E) HBrO₄ Perbromic Acid
F) HNO₂ Nitrous Acid
G) HNO₃ Nitric Acid

- 9) A) tetra phosphorus decaoxide
B) Niobium V oxide
C) Lithium Oxide
D) Titanium IV Nitrate

- 10) A) VF₃
B) O₂F₂
C) Rb₂O₂
D) Ga₂O₃
O₂⁻² = Peroxide