

Chemistry Ch 9
Wkst 9G Net Ionic Reactions

REACTION CATEGORY	NET IONIC REACTIONS
REACTION DESCRIPTION	In describing reactions that occur in solution, it is often desirable to write the equation for the reaction in ionic form, indicating the ionic species that actually exist in solution.
REACTION FORMAT (shown as an example)	$2\text{Na}^{1+}(\text{aq}) + \text{SO}_4^{2-}(\text{aq}) + \text{Ba}^{2+}(\text{aq}) + 2\text{Cl}^{-}(\text{aq}) \rightarrow 2\text{Na}^{1+}(\text{aq}) + 2\text{Cl}^{-}(\text{aq}) + \text{BaSO}_4(\text{s})$
REACTION GUIDELINES	<p>In writing a total ionic equation for reactions in water:</p> <ol style="list-style-type: none"> 1. We indicate all soluble ionic materials as ions, followed by (aq). 2. All substances that react with water to form ions are written as ions followed by (aq). 3. All insoluble ionic solids are written with (s) following their formula. 4. All soluble unionized species are written with their molecular formula followed by (aq). 5. In the above equation, we note that the sodium and chloride ions are unchanged and are present on both sides of the equation. Since they are not undergoing chemical reaction, they can be referred to as spectator ions. If we subtract the spectator ions from each side of the equation, we then have a net ionic equation.

Write the balanced net ionic equation for the following reactions. The starting equations are not necessarily balanced.



