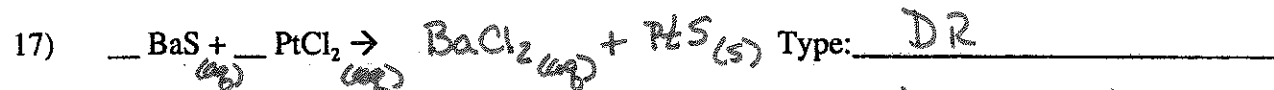
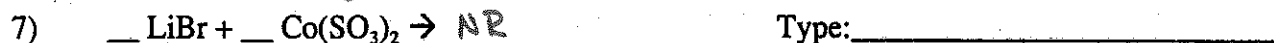
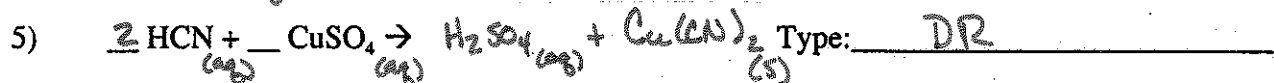


Section 2: Predicting the products of chemical reactions

Predict the products and the type of reaction of the following reactions:



Types of Reactions Worksheet

Section 1: Identify the type of reaction and balancing

Balance the following equations and indicate the type of reaction taking place:

- 1) $\underline{3} \text{NaBr} + \text{H}_3\text{PO}_4 \rightarrow \text{Na}_3\text{PO}_4 + \underline{3} \text{HBr}$ Double Replacement
- 2) $\underline{3} \text{Ca(OH)}_2 + \text{Al}_2(\text{SO}_4)_3 \rightarrow \underline{3} \text{CaSO}_4 + \underline{2} \text{Al(OH)}_3$ DR
- 3) $\underline{3} \text{Mg} + \text{Fe}_2\text{O}_3 \rightarrow \underline{2} \text{Fe} + \underline{3} \text{MgO}$ SR
- 4) $\text{C}_2\text{H}_4 + \underline{3} \text{O}_2 \rightarrow \underline{2} \text{CO}_2 + \underline{2} \text{H}_2\text{O}$ Combustion
- 5) $\underline{2} \text{PbSO}_4 \rightarrow \underline{2} \text{PbSO}_3 + \text{O}_2$ decomposition
- 6) $\underline{2} \text{NH}_3 + \underline{3} \text{I}_2 \rightarrow \text{N}_2\text{I}_6 + \underline{3} \text{H}_2$ Single Replacement
- 7) $\text{H}_2\text{O} + \text{SO}_3 \rightarrow \text{H}_2\text{SO}_4$ Combination
- 8) $\text{Na}_3\text{PO}_4 + \underline{3} \text{KOH} \rightarrow \underline{3} \text{NaOH} + \text{K}_3\text{PO}_4$ DR
- 9) $\text{MgCl}_2 + \text{Li}_2\text{CO}_3 \rightarrow \text{MgCO}_3 + \underline{2} \text{LiCl}$ DR
- 10) $\text{C}_6\text{H}_{12} + \underline{9} \text{O}_2 \rightarrow \underline{6} \text{CO}_2 + \underline{6} \text{H}_2\text{O}$ Combustion
- 11) $\text{Pb} + \text{FeSO}_4 \rightarrow \text{PbSO}_4 + \text{Fe}$ SR
- 12) $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$ Decomposition
- 13) $\text{P}_4 + \underline{3} \text{O}_2 \rightarrow \underline{2} \text{P}_2\text{O}_3$ Combination
- 14) $\underline{2} \text{RbNO}_3 + \text{BeF}_2 \rightarrow \text{Be(NO}_3)_2 + \underline{2} \text{RbF}$ DR
- 15) $\underline{2} \text{AgNO}_3 + \text{Cu} \rightarrow \text{Cu(NO}_3)_2 + \underline{2} \text{Ag}$ SR
- 16) $\text{C}_3\text{H}_6\text{O} + \underline{4} \text{O}_2 \rightarrow \underline{3} \text{CO}_2 + \underline{3} \text{H}_2\text{O}$ Combustion
- 17) $\underline{2} \text{C}_5\text{H}_5 + \text{Fe} \rightarrow \text{Fe(C}_5\text{H}_5)_2$ Combination
- 18) $\text{SeCl}_6 + \text{O}_2 \rightarrow \text{SeO}_2 + \underline{3} \text{Cl}_2$ SR
- 19) $\underline{2} \text{MgI}_2 + \text{Mn(SO}_3)_2 \rightarrow \underline{2} \text{MgSO}_3 + \text{MnI}_4$ DR
- 20) $\text{O}_3 \rightarrow \text{O} + \text{O}_2$ Decomposition
- 21) $\underline{2} \text{NO}_2 \rightarrow \underline{2} \text{O}_2 + \text{N}_2$ Decomposition