

Types of Reactions

Directions: Identify the following reactions using the following: C = combustion S = synthesis
D = decomposition SR = single replacement DR = double replacement
Write Equations + Balance

1. mercury + oxygen -----> mercury (II) oxide
2. zinc + hydrogen sulfate -----> zinc sulfate + hydrogen
3. water -----> hydrogen + oxygen (electrolysis)
4. calcium oxide + water -----> calcium hydroxide
5. sodium chloride + hydrogen sulfate -----> sodium hydrogen sulfate + hydrogen chloride
6. sodium + iodine
7. propane (C₃H₈) + oxygen
8. aluminum + hydrogen sulfate
9. iron (III) sulfate + ammonium sulfide
10. nickel (II) chlorate -----> (heating)
11. copper (II) hydroxide + acetic acid (HC₂H₃O₂) ----->
12. potassium iodide + chlorine
13. zinc hydroxide -----> (heating)
14. zinc + lead (II) acetate ----->
15. hydrogen + oxygen ----->
16. silver nitrate + potassium chloride ----->
17. zinc + hydrogen chloride (HCl) ----->
18. barium carbonate -----> (heating)
19. butene (C₄H₈) + oxygen ----->
20. hydrogen + nitrogen -----> ammonia (NH₃)