

Worksheet 9E -- Chemistry  
Ch. 9 Chemical Reactions

Name Key  
Period \_\_\_\_\_

Double Replacement Reactions

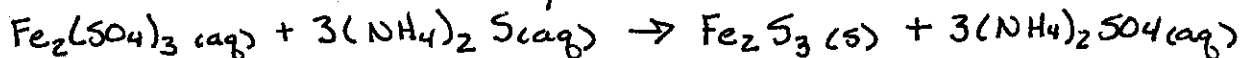
Directions: Complete the word equations. Consult the solubility table on your white sheet to determine if the reaction occurs. Write a complete and balanced equation for each reaction which will occur. If no precipitate forms write N.R. (no reaction). Be sure to include states!

1. silver nitrate plus potassium chloride yields Silver chloride & Potassium Nitrate

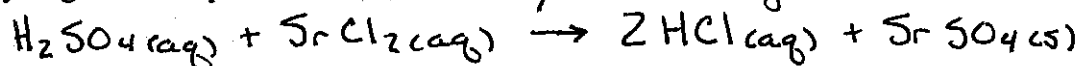


2. sodium chloride plus magnesium nitrate  $\rightarrow$  N.R.

3. iron (III) sulfate plus ammonium sulfide yields Iron III Sulfide & Ammonium Sulfate

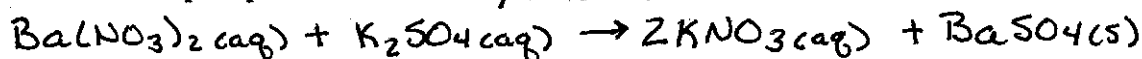


4. hydrogen sulfate plus strontium chloride yields hydrogen chloride & strontium sulfate



5. barium chloride plus sodium hydroxide  $\rightarrow$  N.R.

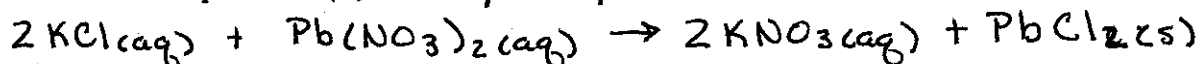
6. barium nitrate plus potassium sulfate yields Barium Sulfate + Potassium Nitrate



7. magnesium carbonate plus hydrogen sulfate yields Magnesium Sulfate + water + Carbon dioxide

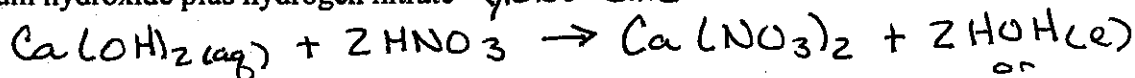


8. potassium chloride plus lead (II) nitrate yields potassium Nitrate & Lead II chloride



9. nickel (II) sulfide plus hydrogen chloride  $\rightarrow$  N.R.

10. calcium hydroxide plus hydrogen nitrate yields Calcium Nitrate and water



or  
 $2\text{H}_2\text{O}(\text{l})$  / 1/04