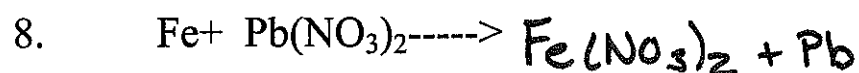
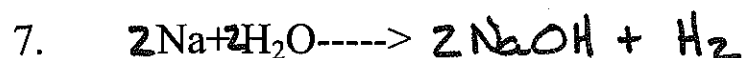
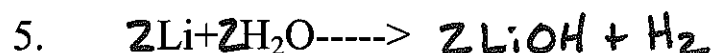
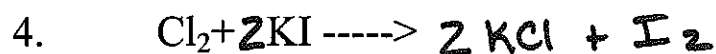
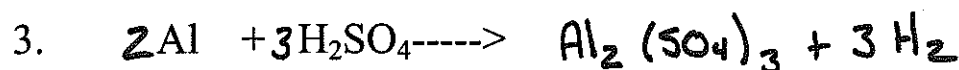
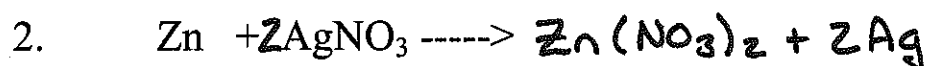


REACTION CATEGORY	SINGLE REPLACEMENT REACTION
REACTION DESCRIPTION	<p>In these reactions, a free element reacts with a compound to form another compound and release one of the elements of the original compound in the elemental state. There are two different possibilities:</p> <ol style="list-style-type: none"> 1. One cation (+ ion) replaces another. 2. One anion (- ion) replaces another.
REACTION FORMAT	<ol style="list-style-type: none"> 1. $AB + C \rightarrow CB + A$ 2. $A + BC \rightarrow BA + C$
<p>REACTION GUIDELINES</p> <p>Assume Iron II in Eqns !!</p>	<ol style="list-style-type: none"> 1. In a single replacement reaction atoms of one element replace the atoms of a second element in a compound. Whether one metal will replace another metal from a compound can be determined by the relative reactivities of the two metals. To help us determine this, an activity series of metals arranges metals in order of decreasing reactivity. A reactive metal will replace any metal listed below it in the activity series. 2. A nonmetal can also replace another nonmetal from a compound. This replacement is usually limited to the halogens (F_2, Cl_2, Br_2, and I_2). The activity of the halogens decreases as you go down the Group (17) of the periodic table.



Chemistry Ch 9

Wkst 9D2 Single Replacement

