

## Worksheet: Understanding Formulas

(Frameworks Code)

Part A: Complete the following table:

	Name of Positive Ion	Name of Negative Ion	Name of Compound
1. KOH	Potassium	hydroxide	Potassium hydroxide
2. NaPO <sub>4</sub>	Sodium	Phosphate	Sodium Phosphate
3. Ca(ClO <sub>3</sub> ) <sub>2</sub>	Calcium	chlorate	Calcium chlorate
4. MgSO <sub>4</sub>	Magnesium	Sulfate	magnesium Sulfate
5. NH <sub>4</sub> HCO <sub>3</sub>	Ammonium	Bicarbonate	Ammonium Bicarbonate
6. LiNO <sub>3</sub>	Lithium	Nitrate	Lithium Nitrate
7. BaCO <sub>3</sub>	Barium	Carbonate	Barium Carbonate
8. KC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>	Potassium	Acetate	Potassium Acetate
9. NH <sub>4</sub> OH	Ammonium	Hydroxide	Ammonium Hydroxide

Part B: Complete the following table:

Formula	Symbol & oxidation # of the + Ion	Symbol & oxidation # of the - Ion	Formula	Symbol & oxidation # of the + Ion	Symbol & oxidation # of the - Ion
Ex: AlPO <sub>4</sub>	Al 3+	PO <sub>4</sub> 3-	8. Ba <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub>	Ba <sup>+2</sup>	PO <sub>4</sub> <sup>-3</sup>
1. CaCl <sub>2</sub>	Ca <sup>+2</sup>	Cl <sup>-</sup>	9. Cu <sub>2</sub> O	Cu <sup>+1</sup>	O <sup>-2</sup>
2. Ag <sub>2</sub> S	Ag <sup>+1</sup>	S <sup>-2</sup>	10. HgCl <sub>2</sub>	Hg <sup>+2</sup>	Cl <sup>-</sup>
3. Pb(CO <sub>3</sub> ) <sub>2</sub>	Pb <sup>+4</sup>	CO <sub>3</sub> <sup>-2</sup>	11. PbO	Pb <sup>+2</sup>	O <sup>-2</sup>
4. K <sub>2</sub> SO <sub>4</sub>	K <sup>+</sup>	SO <sub>4</sub> <sup>-2</sup>	12. Fe <sub>2</sub> O <sub>3</sub>	Fe <sup>+3</sup>	O <sup>-2</sup>
5. HgCl	Hg <sup>+</sup>	Cl <sup>-</sup>	13. FeS	Fe <sup>+2</sup>	S <sup>-2</sup>
6. SnO <sub>2</sub>	Sn <sup>+4</sup>	O <sup>-2</sup>	14. H <sub>2</sub> S	H <sup>+</sup>	S <sup>-2</sup>
7. PbCl <sub>2</sub>	Pb <sup>+2</sup>	Cl <sup>-</sup>	15. NH <sub>4</sub> OH	NH <sub>4</sub> <sup>+1</sup>	OH <sup>-</sup>