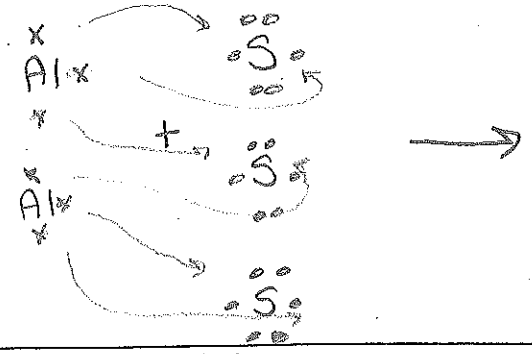
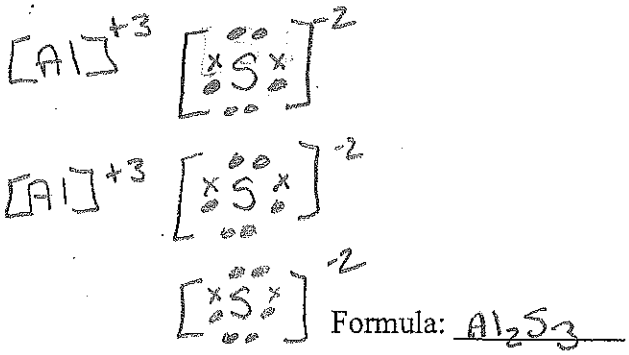
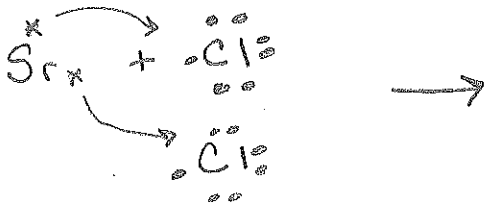
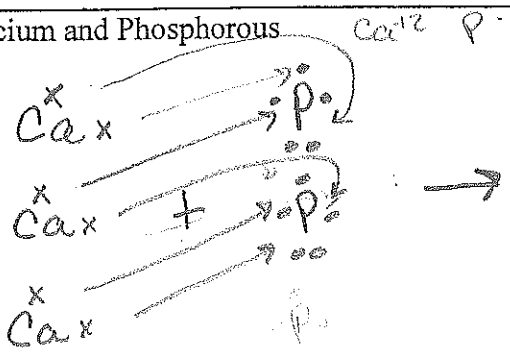
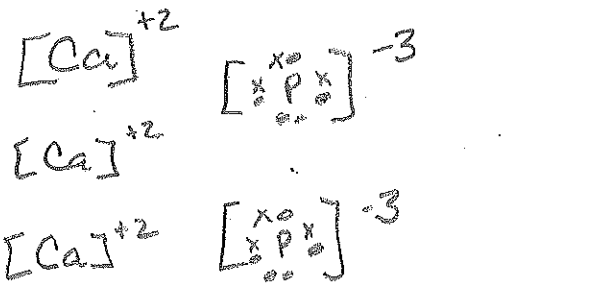
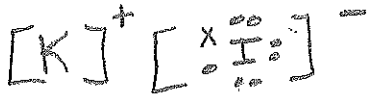
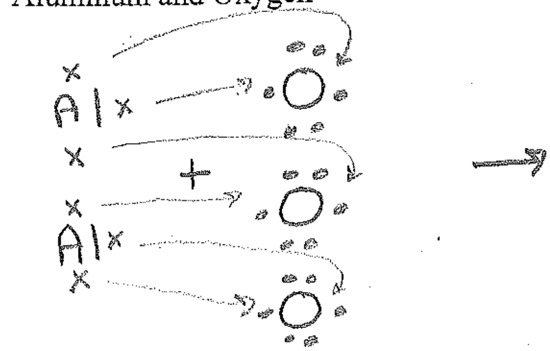
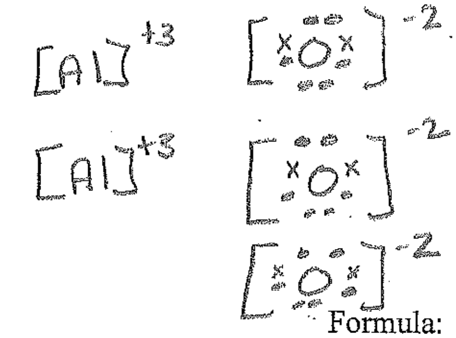
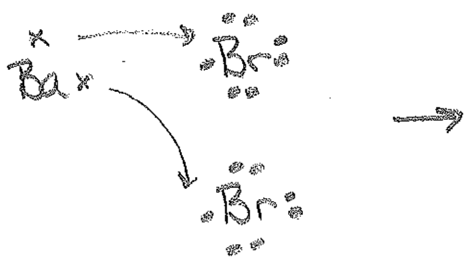
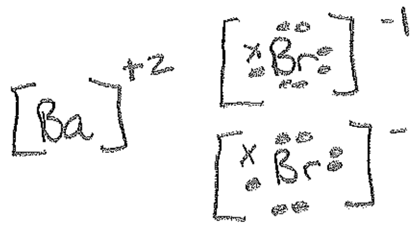
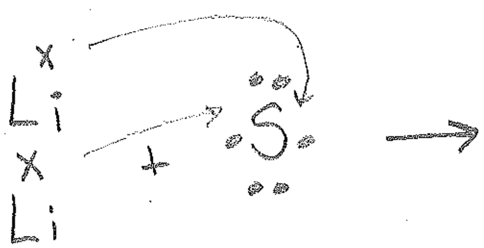
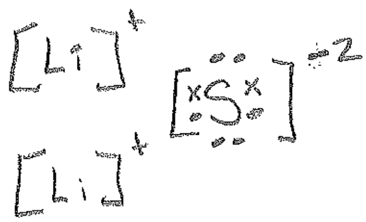
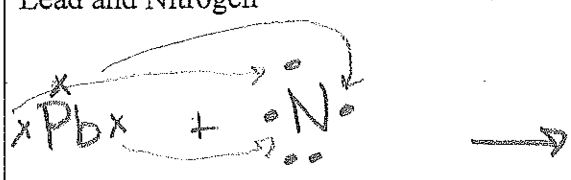
WKST 20.B

Ionic Compound Drawing

Directions:

1. Using electron dot diagrams and arrows, indicate how the following elements will transfer electrons to form stable ionic bonds
2. Then use brackets and + or – signs to indicate the charge of the resulting ions.
3. Finally, write the formula for the resulting compound in the blank provided.

Aluminum and Sulfur 	 Formula: Al_2S_3
Strontium and Chlorine 	 Formula: $SrCl_2$
Calcium and Phosphorous 	 Formula: Ca_3P_2
Potassium and Iodine 	 Formula: KI

Aluminum and Oxygen 	 Formula: Al_2O_3
Barium and Bromine 	 Formula: $BaBr_2$
Lithium and Sulfur 	 Formula: Li_2S
Lead III and Nitrogen 	 Formula: PbN