

UNDERSTANDING IONS

Element	Group Number	# of Valence Electrons	Oxidation Number	Number of electrons gained or lost	Ionic Form	Cation or Anion
Mg	2	2	+2	Will Lose 2	Mg ²⁺	Cation (+)
Na	1	1	+1	will lose 1	Na ⁺	Cation ⁺
Ca	2	2	+2	will lose 2	Ca ⁺²	Cation
Al	13 (3)	3	+3	will lose 3	Al ⁺³	Cation
P	15 (5)	5	-3	will gain 3	P ⁻³	Anion
S	16 (6)	6	-2	will gain 2	S ⁻²	Anion
Br	17 (7)	7	-1	will gain 1	Br ⁻¹	Anion
Li	1	1	+1	will lose 1	Li ⁺	Cation
Cl	17 (7)	7	-1	will gain 1	Cl ⁻	Anion
Zn (II)	12	2	+2	will lose 2	Zn ⁺²	Cation
Fe (III)	8	3	+3	will lose 3	Fe ⁺³	Cation
Ba	2	2	+2	will lose 2	Ba ⁺²	Cation
F	17 (7)	7	-1	will gain 1	F ⁻¹	Anion
K	1	1	+1	will lose 1	K ⁺	Cation
I	17 (7)	7	-1	will gain 1	I ⁻¹	Anion
Cs	1	1	+1	will lose 1	Cs ⁺¹	Cation
N	15 (5)	5	-3	will gain 3	N ⁻³	Anion
Ag (I)	11	1	+1	will lose 1	Ag ⁺	Cation
O	16 (6)	6	-2	will gain 2	O ⁻²	Anion
Cu (II)	11	2	+2	will lose 2	Cu ⁺²	Cation
Cr (III)	6	3	+3	will lose 3	Cr ⁺³	Cation