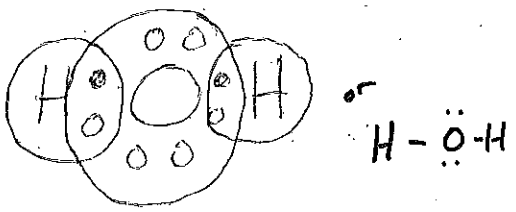
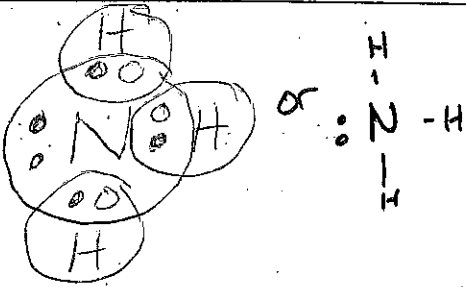
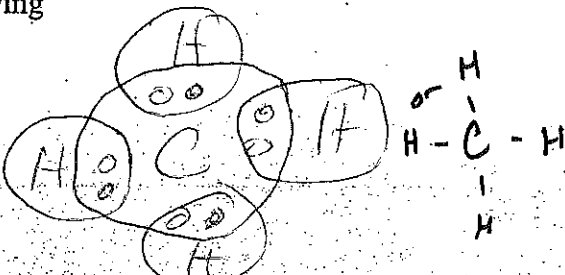
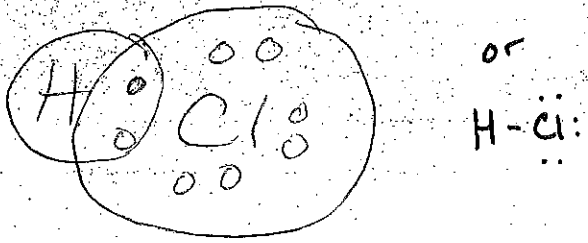
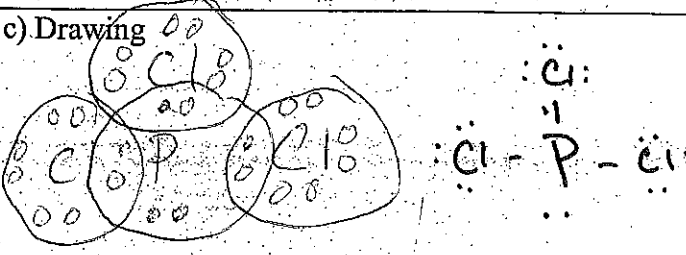
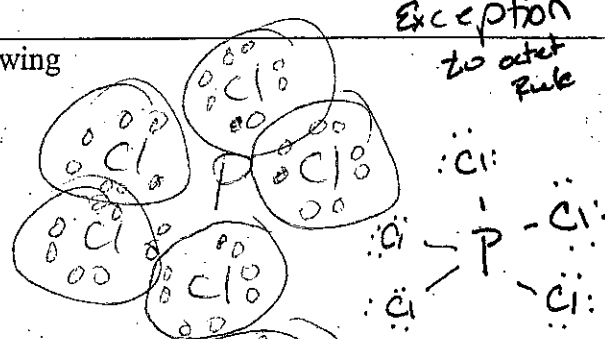
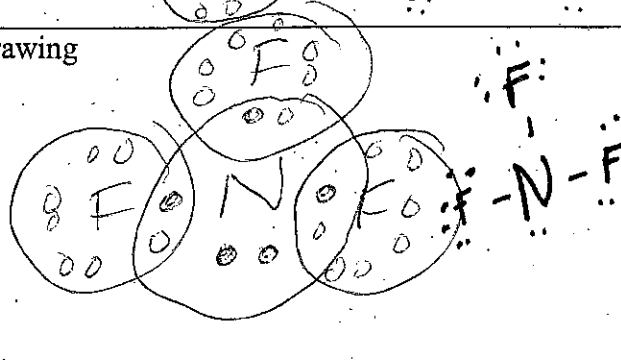
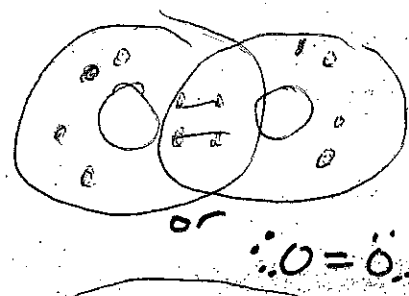
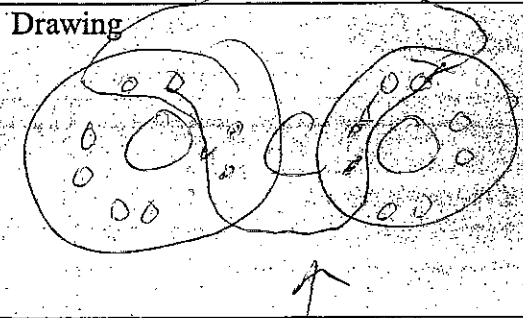


Name Key

Drawing Covalent Compounds

<p>1. H_2O</p> <p>a) Total # of outer electrons: $H-2$ $O-8$</p> <p>b) Compound Name: <i>Dihydrogen Monoxide</i></p>	<p>c) Drawing</p> 
<p>2. NH_3</p> <p>a) Total # of outer electrons: $N-8$ $O-2$</p> <p>b) Compound Name: <i>Nitrogen trihydride</i></p>	<p>c) Drawing</p> 
<p>3. CH_4</p> <p>a) Total # of outer electrons: $H-2$ $O-8$</p> <p>b) Compound Name: <i>Carbon tetrachloride</i></p>	<p>c) Drawing</p> 
<p>4. HCl</p> <p>a) Total # of outer electrons: $H-2$ $Cl-8$</p> <p>b) Compound Name: <i>Hydrogen Monochloride</i></p>	<p>c) Drawing</p> 
<p>5. PCl_3</p> <p>a) Total # of outer electrons: $Cl-8$ $P-8$</p> <p>b) Compound Name: <i>Carbon trichloride</i></p>	<p>c) Drawing</p> 

<p>6. PCl_5</p> <p>a) Total # of outer electrons: $\text{P} - 5$ $\text{Cl} - 7$</p> <p>b) Compound Name: <i>Phosphorus pentachloride</i></p>	<p>c) Drawing</p>  <p><i>P Exception to octet Rule</i></p>
<p>7. NF_3</p> <p>a) Total # of outer electrons: $\text{N} - 5$ $\text{F} - 7$</p> <p>b) Compound Name: <i>Nitrogen trifluoride</i></p>	<p>c) Drawing</p> 
<p>8. O_2</p> <p>a) total # of outer electrons: $\text{O} - 6$ <i>Double Bond</i></p> <p>b) Compound Name: <i>Diatomic Oxygen</i></p>	<p>c) Drawing</p>  <p>or $:\text{O}=\text{O}:$</p>
<p>9. CO_2</p> <p>a) total # of outer electrons: $\text{C} - 4$ $\text{O} - 6$ <i>Double Bond</i></p> <p>b) Compound Name: <i>Carbon Dioxide</i></p>	<p>c) Drawing</p>  <p>or $:\text{O}=\text{C}=\text{O}:$</p>