

# Ch 18 Pretest - Answers

- 1) Sulfuric Acid
- 2) Sulfurous Acid
- 3)  $\text{HNO}_3$
- 4)  $\text{Ca}(\text{OH})_2$
- 5)

$$\text{pOH} = -\log[\text{OH}^-]$$

$$= -\log[5.34 \times 10^{-10}]$$

$$\boxed{\text{pOH} = 9.27}$$

$$\text{pH} = 14 - \text{pOH}$$

$$= 14 - 9.27$$

$$\boxed{\text{pH} = 4.73}$$

$$6) \left( \frac{5.0 \text{ g HBr}}{1} \right) \left( \frac{1 \text{ mole HBr}}{80.9 \text{ g HBr}} \right) \left( \frac{1}{.45 \text{ L}} \right) = .14 \text{ M HBr}$$

$$\text{pH} = -\log[\text{H}^+]$$

$$= -\log[.14]$$

$$\boxed{\text{pH} = .85}$$

$$\text{pOH} = 14 - \text{pH}$$

$$= 14 - .85$$

$$\boxed{\text{pOH} = 13.15}$$

$$7) \text{pOH} = -\log[\text{OH}^-]$$

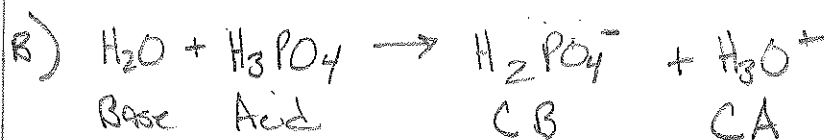
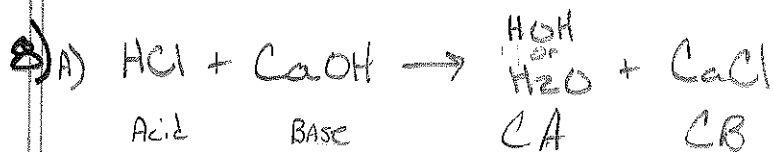
$$= -\log[6.3 \times 10^{-9}]$$

$$\boxed{\text{pOH} = 8.2}$$

$$\text{pH} = 14 - \text{pOH}$$

$$= 14 - 8.2$$

$$\boxed{\text{pH} = 5.8}$$



9) pH stands for "potential of hydrogen"

10) Diprotic Acids can donate more than 1 hydrogen into solution  
Example  $\text{H}_2\text{SO}_4$

11) A substance that can act as both an Acid & Base