

# Can you make 2.00 grams of a Compound?

## Chemistry Chapter 11 - Stoichiometry

### Introduction

Use your skills of predicting chemical reactions, balancing equations, and calculating molar mass to solve a complex stoichiometry problem. Then test your lab techniques by mixing the reactants and isolating exactly 2.00 g of a compound.

### Step 1: Draw your reactants out of a beaker

#### Prelab :

1. Write a balanced equation for the reaction that will take place with the two reactants you have been picked.
2. Predict which of the products formed will form a precipitate using your solubility table.
3. Calculate the mass of the reactants needed to produce the 2.00 grams of precipitate you will make in the lab.

### Lab Procedure

**Purpose:** Use the calculated mass of reactants to make and recover 2.00 grams of precipitate.

#### Procedure:

1. Both reactants need to be dissolved in separate beakers using 25 mL of distilled water. When each of the reactants have dissolved, mix the two solutions together to form a precipitate.
2. Recover the precipitate by filtration using a pre-massed piece of filter paper. Using proper setup for filtering.
3. Filter paper with precipitate will be dried and then the mass determined.
4. Show Mr Allan your Final mass on a scale

#### Data & results:

- Data table to include
- (1) Mass of filter paper
  - (2) mass of filter paper and ppt after drying
  - (3) Mass of ppt

#### Conclusion:

Calculate % Error and explain where your percent error occurred from. (Complete sentences)